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$I H P_{vap} = \exp 14.790 - T K$ and by the ideal gas law f Solutions to Chemical and Engineering Thermodynamics, $3e N V RT \times 18 \text{ mol} \times 8.314 \times 10^{-5} \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{K}^{-1})$ $T P = \frac{V V_1 \times 10^{-3} - (10 - x) 106}{\text{volume taken up by liquid}}$ Equating P and P_{vap} we have $x \times 18 \times 8.314 \times 10^{-5} T F 5432.8 \times I$.

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